NAME:\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ DATE:\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

**Chemistry Periodic Table (Chapter 6 Outline)**

***AS YOU READ, RESPOND TO THE FOLLOWING:***

*Vocabulary*: **Define in your own words**

Periodic Law-

Group-

Period-

Representative element-

Transition element-

Halogen-

Noble Gas-

Metalloid-

Ion-

Ionization energy-

Electronegativity-

Valence electrons-

Respond to the prompts below:

What are the key features of the Periodic Table? Explain. (3 pts)

Why do elements in the same group have similar properties? Explain. (2 pts)

How are period and group trends in atomic radii related to electron configuration? Explain. (3 pts)

***ABOUT THE READING:***

Write three things that you learned about the Periodic Table*:*

***Make sure to write a full sentence.***

*Example: I learned that John Newlans proposed the "Law of Octaves", which asserted that there was a periodic relationship between properties in every eighth element when elements are arranged according to atomic mass.*

*1.*

*2.*

*3.*

***Assigned work****: Chapter 6 assessment, problems #25, 27, 31, 35, 41, 47, 48, 54-56, 61-65*